

Density L2 Mark Scheme		
<b>1(a)</b>	$1 \div 1 (= 1)$	[1]
	$1 \text{ g/cm}^3$	[1]
<b>1(b)</b>	$1 \times 1 (= 1)$	[1]
	$1 \text{ g}$	[1]
<b>1(c)</b>	$1 \div 1 (= 1)$	[1]
	$1 \text{ cm}^3$	[1]
<b>1(d)</b>	$1 \div 1000 (= 0.001)$	[1]
	$0.001 \text{ m}^3$	[1]
<b>1(e)</b>	$5 \times 150 (= 750)$	[1]
	$750 \text{ g}$	[1]
<b>1(f)</b>	$8300 \div 16600 (= 0.5)$	[1]
	$0.5 \text{ g/cm}^3$	[1]
<b>1(g)</b>	$5040 \div 21 (= 240)$	[1]
	$240 \text{ cm}^3$	[1]
<b>1(h)</b>	$756 \div 252 (= 3)$	[1]
	$3 \text{ g/cm}^3$	[1]
<b>1(i)</b>	$0.00125 \times 100000 (= 125)$	[1]
	$125 \text{ g}$	[1]
<b>1(j)</b>	$65536 \div 0.00390625 (= 16777216)$	[1]
	$16777216 \text{ kg/m}^3$	[1]
<b>2(a)</b>	$500 \div 500 (= 1)$	[1]
	$= 1 \text{ g/cm}^3$	[1]
<b>2(b)</b>	$625 \div 500 (= 1.25)$	[1]
	$1.25 \text{ g/cm}^3$	[1]
<b>2(c)</b>	$375 \div 125 (= 3)$	[1]
	$3 \text{ g/cm}^3$	[1]
<b>2(d)</b>	$300 \div 250 (= 1.2)$	[1]
	$1.2 \text{ g/cm}^3$	[1]

<b>3(a)</b>	$7.9 \times 8 (= 63.2)$	[1]
	63.2 g	[1]
<b>3(b)</b>	$2.7 \times 64 (= 172.8)$	[1]
	172.8 g	[1]
<b>3(c)</b>	$11.3 \times 27 (= 305.1)$	[1]
	305.1 g	[1]
<b>3(d)</b>	$2.2 \times 216 (= 475.2)$	[1]
	$3.0 \times 125 (= 375)$	[1]
	Sandstone has the greatest mass because 475.2 is greater than 375	[1]
<b>4(a)</b>	$200 \div 1 (= 200)$	[1]
	A: 200 cm <sup>3</sup>	[1]
	$1320 \div 1.1 (= 1200)$	[1]
	B: 1200 cm <sup>3</sup>	[1]
<b>4(b)</b>	$800 \div 1.6 (= 500)$	[1]
	$500 \times 2 (= 1000)$	[1]
	1000 cm <sup>3</sup>	[1]
<b>4(c)</b>	D: $10201 \div 1.01 (= 10100)$	[1]
	E: $10000 \div 1.25 (= 8000)$	[1]
	D is larger.	[1]
	$10100 - 8000 = 2100 \text{ cm}^3$	[1]
<b>5(a)</b>	$2000 \div 8 (= 250)$	[1]
	250 kg/m <sup>3</sup>	[1]
<b>5(b)</b>	$5000 \div 25 (= 200)$	[1]
	200 kg/m <sup>3</sup>	[1]
	No	[1]
<b>5(c)</b>	$2500 \div 6.25 (= 400)$	[1]
	$1200 \div 4 (= 300)$	[1]
	Denser than Tayata Riva, Valvo Bus and Fait 220	[1]

<b>6(a)</b>	Lead volume: $1 \times 1 \times 1 = 1 \text{ cm}^3$	[1]
	Aluminium volume: $8 \times 8 \times 8 = 512 \text{ cm}^3$	[1]
	Iron volume: $4 \times 4 \times 4 = 64 \text{ cm}^3$	[1]
	Gold volume: $3 \times 3 \times 3 = 27 \text{ cm}^3$	[1]
	Lead density: $13824 \div 1 = 13824 \text{ g/cm}^3$	[1]
	Aluminium density: $13824 \div 512 = 27 \text{ g/cm}^3$	[1]
	Iron density: $13824 \div 64 = 216 \text{ g/cm}^3$	[1]
	Gold density: $13824 \div 27 = 512 \text{ g/cm}^3$	[1]
<b>6(b)</b>	$1 + 512 = 513 \text{ cm}^3$	[1]
	$5130 \div 513$	[1]
	$= 10 \text{ g/cm}^3$	[1]